

*****LAB SAMPLE DATA*****

The Decomposition of Potassium Chlorate

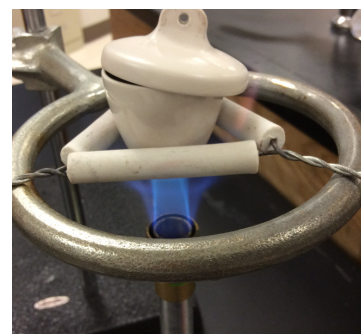
1. Determine the mass of the empty crucible.



2. Add approximately two grams KClO_3 to the crucible.



3. Heat the crucible over a Bunsen burner for at least 15 minutes (until all the powder has reacted.)



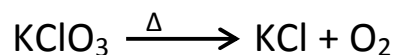
4. And we wait.....

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5. Let the crucible cool AT LEAST 5 minutes....
6. Reweigh the crucible with the heated contents.



This was a decomposition reaction as follows:



The mass at the end of the reaction is less than the mass at the beginning of the reaction because the oxygen escaped as a gas and only the potassium chloride remained in the crucible.

You should be able to use this sample data to complete the lab and the calculations that are required. Please attend tutorials if you need help.