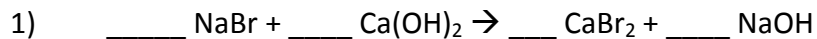


Mole Ratios

Name _____ Period _____

Balance the following chemical reactions and identify the mole ratios.



What type of reaction: _____

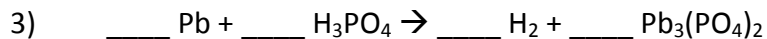
Mole ratios:

$\frac{\text{mol NaBr}}{\text{mol Ca(OH)}_2}$	$\frac{\text{mol CaBr}_2}{\text{mol NaBr}}$	$\frac{\text{mol NaBr}}{\text{mol NaOH}}$	$\frac{\text{mol CaBr}_2}{\text{mol Ca(OH)}_2}$
$\frac{\text{mol Ca(OH)}_2}{\text{mol NaBr}}$	$\frac{\text{mol NaBr}}{\text{mol CaBr}_2}$	$\frac{\text{mol NaOH}}{\text{mol NaBr}}$	$\frac{\text{mol Ca(OH)}_2}{\text{mol CaBr}_2}$
$\frac{\text{mol Ca(OH)}_2}{\text{mol NaOH}}$	$\frac{\text{mol NaOH}}{\text{mol Ca(OH)}_2}$	$\frac{\text{mol CaBr}_2}{\text{mol NaOH}}$	$\frac{\text{mol NaOH}}{\text{mol CaBr}_2}$



What type of reaction: _____

Mole ratios: (There should be 6 ratios)



What type of reaction: _____

Mole ratios:

- What are the mole ratios between lead and lead (II) phosphate?
- What are the mole ratios hydrogen and phosphoric acid?
- What is the mole ratio between phosphoric acid and lead (II) phosphate?

Answer the following questions. Make sure you balance the equation FIRST.

4) Given this equation: $\text{___ N}_2 + \text{___ H}_2 \rightarrow \text{___ NH}_3$, write the following molar ratios:

- N_2 / H_2
- N_2 / NH_3
- H_2 / NH_3

5) Given the following equation: $\text{___ H}_2 + \text{___ S}_8 \rightarrow \text{___ H}_2\text{S}$, write the following molar ratios:

- $\text{H}_2 / \text{H}_2\text{S}$
- H_2 / S_8
- $\text{H}_2\text{S} / \text{S}_8$

6) Answer the following questions for this equation: $\text{___ H}_2 + \text{___ O}_2 \rightarrow \text{___ H}_2\text{O}$

- What is the $\text{H}_2 / \text{H}_2\text{O}$ molar ratio?
- If you had 20.0 moles of H_2 on hand and plenty of O_2 , how many moles of H_2O could you make?
- What is the $\text{O}_2 / \text{H}_2\text{O}$ molar ratio?
- Suppose you had 20.0 moles of O_2 and enough H_2 , how many moles of H_2O could you make?

7) Use this equation: $\text{___ N}_2 + \text{___ H}_2 \rightarrow \text{___ NH}_3$, for the following problems:

- If you used 1.0 mole of N_2 , how many moles of NH_3 could be produced?
- If 10.0 moles of NH_3 were produced, how many moles of N_2 would be required?
- If 3.00 moles of H_2 were used, how many moles of NH_3 would be made?
- If 0.600 moles of NH_3 were produced, how many moles of H_2 are required?