Writing Formulas Name Key Mr. Sudbury________ Period____Date____

Write the formula in the box that is a result of the row intersecting the column. Remember that a polyatomic ion must stay intact and go in parentheses if a number gets criss-crossed down to it. You also need to reduce if you can.

	Cl^{1-}	O^{2-}	N^{3-}	OH^{1-}	NO_3^{1-}	CO_3^{2-}	SO_4^{2-}	PO_4^{3-}
H^{+}	HCl	H_2O	H_3N	$H(OH)$ or H_2O	HNO_3	H_2CO_3	H_2SO_4	H_3PO_4
Na^{+}	$NaCl$	Na_2O	Na_3N	$NaOH$	$NaNO_3$	Na_2CO_3	Na_2SO_4	Na_3PO_4
Mg^{2+}	$MgCl_2$	reduced MgO	Mg_3N_2	$Mg(OH)_2$	$Mg(NO_3)_2$	reduced $MgCO_3$	reduced $MgSO_4$	$Mg_3(PO_4)_2$
K^{+}	KCl	K_2O	K_3N_2	KOH	KNO_3	K_2CO_3	K_2SO_4	K_3PO_4
Al^{3+}	$AlCl_3$	Al_2O_3	AlN reduced	$Al(OH)_3$	$Al(NO_3)_3$	$Al_2(CO_3)_3$	$Al_2(SO_4)_3$	$AlPO_4$ reduced
Ca^{2+}	$CaCl_2$	CaO reduced	Ca_3N_2	$Ca(OH)_2$	$Ca(NO_3)_2$	$CaCO_3$ reduced	$CaSO_4$ reduced	$Ca_3(PO_4)_2$
NH_4^{+}	NH_4Cl	$(NH_4)_2O$	$(NH_4)_3N$	NH_4OH	NH_4NO_3	$(NH_4)_2CO_3$	$(NH_4)_2SO_4$	$(NH_4)_3PO_4$
Pb^{2+}	$PbCl_2$	reduced PbO	Pb_3N_2	$Pb(OH)_2$	$Pb(NO_3)_2$	reduced $PbCO_3$	reduced $PbSO_4$	$Pb_3(PO_4)_2$
Pb^{4+}	$PbCl_4$	PbO_2 reduced	Pb_3N_4	$Pb(OH)_4$	$Pb(NO_3)_4$	$Pb(CO_3)_2$ reduced	reduced $Pb(SO_4)_2$	$Pb_3(SO_4)_4$
Fe^{2+}	$FeCl_2$	reduced FeO	Fe_3N_2	$Fe(OH)_2$	$Fe(NO_3)_2$	$FeCO_3$	reduced $FeSO_4$	$Fe_3(PO_4)_2$
Fe^{3+}	$FeCl_3$	Fe_2O_3	FeN reduced	$Fe(OH)_3$	$Fe(NO_3)_3$	$Fe_2(CO_3)_3$	$Fe_2(SO_4)_3$	$FePO_4$ reduced